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with Moseley ' s
revision, the Periodic
Law states: the
properties of the
elements are periodic
functions of their
atomic numbers.
Each of the 118

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known elements has its own set of characteristic properties. The elements are arranged within the periodic table into groups and periods.

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Purpose: In this
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Use your knowledge of periodic properties and a list of clues to correctly arrange the elements from a scrambled periodic table. You will also predict values for any information missing from the table.

Theory: The present organization of the elements is a product of the first periodic

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table published by
Dmitri Mendeleev in
1869.

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Periodic Law is considered to be one of the most important concepts in chemistry. Every chemist makes use of Periodic Law, whether consciously or not, when dealing with the chemical elements, their properties, and their chemical reactions.

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Periodic Law led to the development of the modern periodic table. Discovery of Periodic Law

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periodic law was
developed
independently by
Dmitri Mendeleev
and Lothar Meyer in
1869. Mendeleev

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created the first periodic table and was shortly followed by Meyer. They both arranged the elements by their mass and proposed that certain properties periodically reoccur. Meyer formed his periodic law based on the atomic volume or molar volume, which

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is the atomic mass
divided by the
density in solid form.

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periodic table. 2.
Explain why elements
in the same group
have similar

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properties. 3. Identify the four blocks of the periodic table based on electron

configuration. 4.

Compare period and group trends of several properties. 5.

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P is less dense than S.
S is an alkali metal. E

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is a noble gas. In this activity, students generate a periodic table from clues and predict the missing properties of several elements based on the elements' locations in the table.

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