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EXAMPLES 1A (E) A has a mass number of zero and an "atomic number" of 1. Emission of this electron has the effect of transforming a neutron into a proton. 0_0e 94 ${}^{241}_{95}\text{Pu}$ Am ${}^{+0}_1\text{B}$ (E) 58 Ni has a mass number of 58 and an atomic number of 28. A positron

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has a mass

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Chapter 25: Nuclear Chemistry 1197 (c) We now determine the number of atoms remaining after 16 days. Because two half-lives elapse in 16 days, the number of atoms has been halved twice, to one-fourth (25%) the original number of atoms.

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answers. 31. Implementation of Texas Essential
Knowledge and Skills for Science, High School. (a) The
provisions of this subchapter shall be implemented by
school districts Chapter 25 nuclear chemistry practice
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Chapter 20 SG 20.1 □ Determining Oxidation Numbers
□ SG 20.2 & 20.3 What is Oxidizing and Reducing?
Balancing with Oxidation Numbers Using Half
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